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the left. Show all your work in the space provided. 1.
_____ The actual yield of a reaction is 22 g and the
theoretical yield is 25 g. Calculate the percentage
yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of
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equation above, how many moles of NH_3 are produced
from each mole of N_2 ? ... modern chemistry 80
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Stoichiometry 1. Given the following equation: $\text{C}_3\text{H}_4(\text{g})$
 $+ 4\text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$ or a. What is the mole
ratio of O_2 to 1-120 in the above equation? b ... ISD
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Date | Chapter Test B, continued PART III Write the
answers to the questions on the line to the left, and
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